Additions and Corrections

NOTICE TO READERS.—For the convenience of those who may wish to cut out the corrections and attach them to the margins of the articles corrected, they have been printed only upon one side of the page.

1930, VOL. 52

Rotenone. II. The Derivatives of Derritol. By F. B. LaForge and L. E. Smith.

Page 1089. In lines 37–38, read "... The theoretical volume of hydrogen corresponding to the addition of one molecule is 58.4 cc. per gram...."

Page 1090. In line 18 for "$2CH_3O$" read "$3CH_3O$." In lines 21–22 read "... The theoretical volume of hydrogen corresponding to the addition of one molecule is 56.6 cc. per gram...."—F. B. LAFORGE.

1932, VOL. 54

The Interaction of Diarylarsyl Iodides, Diarylstibyl Iodides and Phenyldihaloarsines with the Piperidine Salt of N-Pentamethylenedithiocarbamic Acid. By F. F. Blicke and U. O. Oakdale.

Page 2996. In line two the melting point of phenylarsylene N-pentamethylenedithiocarbamate is "183–184°" instead of "173–174°."—F. F. BLICKE.

The Activity Coefficients of Potassium Chloride. An Application of the Extended Debye–Hückel Theory to Interpretation of Freezing Point Measurements. By Hugh M. Spencer.

Pages 4495 and 4496. Replace columns in Tables III and IV by corrected data as shown.

	TABLE III, p. 4495			TABLE IV, p. 4496	
	Col. 4	Col. 6	Col. 7	Col. 11	Col. 12
m	$-\frac{2}{2.3026}\int_0^m \frac{j}{m^{1/2}}\,dm^{1/2}$	$-\log \gamma'$	γ'	$-\log \gamma$	γ
0	0	0	1		
0.001	−0.01000	0.01489	0.9663		
.002	− .01402	.02080	.9532		
.005	− .02174	.03204	.9289		
.01	− .03009	.04399	.9037	0.04405	0.9035
.02	− .04126	.05971	.8715	.05988	.8712
.05	− .06145	.08723	.8181	.08767	.8172
.1	− .08176	.11364	.7698	.11390	.7693
.2	− .10598	.14405	.7178	.14252	.7203
.35	− .12895	.17214	.6728	.16760	.6798
.5	− .14517	.19166	.6432	.18384	.6549
.75	− .16504	.21490	.6097	.20228	.6277
1.0	− .18005	.23214	.5860	.21369	.6114
1.25	− .19215	.24542	.5683	.22135	.6007
1.50	− .20225	.25593	.5547	.22645	.5937

HUGH M. SPENCER

1933, VOL. 55

Extra Specific Heat in Cuprous Sulfide; Specific Heat of Ferrous Oxide. By Walter P. White.

Pages 1047 ff.: "In this article were published data on the specific heat up to 900° of a sample of FeO containing a slight admixture of magnetite. The pieces of this sample were analyzed microscopically for magnetite by Dr. H. E. Merwin, and a plausible correction was applied giving the specific heat of the FeO alone. Chemical analysis, which would destroy the material, was deferred until it should be certain that no more specific heat observations were to be made. That time has now been reached, and Dr. J. F. Schairer has very kindly analyzed the whole material. Three analyses by two different methods, agreeing to the probable precision of a determination, 0.01%, give 88.31% FeO [11.69%] Fe_2O_3. The number of atoms per gram in this sample is to that in simple FeO as 1.0146:1. In view of the smallness of the correction it may be assumed that the mean atomic heat of FeO is the same as that of the sample. The heat of FeO is then 0.9856 of that observed, or

interval specific heat is
$0.17256 + 0.0000133t$
the true specific heat, therefore,
$0.17256 + 0.0000266t$."

WALTER P. WHITE

The Chemical Action of Audible Sound. By Earl W. Flosdorf and Leslie A. Chambers.

Page 3051. Text line 31 should read "reduction of the metallic ion by nickel will also take place...."—EARL W. FLOSDORF AND LESLIE A. CHAMBERS.

Tetraarylarsonium Halides. By F. F. Blicke and C. Marzano.

Page 3056. Near the end of the second paragraph, the melting point of tetraphenylarsonium chloride is "256–257°" instead of "272–274°"; that of the arsonium iodide "312–313°" instead of "292–293°."—F. F. BLICKE.

The Preparation of Certain Cryptophenols. By Larkin Hundley Farinholt, Wilton C. Harden and Daniel Twiss.

Page 3386. In the second part of the table at the top of the page the boiling point of 2,4-dibenzylphenol should be "248 (10 mm.)."—L. H. FARINHOLT.

The Partial and Integral Heats of Dilution of Cadmium Sulfate Solutions from Electromotive Force Measurements. By Victor K. La Mer and W. George Parks.

Page 4350. The authors write: "In making the calculations of Tables VI and VII, the original data for 0.01 m $CdSO_4$ were used instead of the smoothed values. The correct values for the coefficients in Equation 14 for 0.01 m $CdSO_4$ should be

$A = 3812$
$B = -19.5333$
$C = -1.47$
$D = 0.015333$